

Standard Reduction Potentials @ 25°C

Half Reaction	Potential (E°)	Half Reaction	Potential (E°)
$F_2 + 2e^- \rightarrow 2F^-$	+2.87 V	$2H^+ + 2e^- \rightarrow H_2$	0.00 V
$O_3 + 2H^+ + 2e^- \rightarrow O_2 + H_2O$	+2.07 V	$Fe^{3+} + 3e^- \rightarrow Fe$	-0.04 V
$S_2O_8^{2-} + 2e^- \rightarrow 2SO_4^{2-}$	+2.05 V	$Pb^{2+} + 2e^- \rightarrow Pb$	-0.13 V
$PbO_2 + 4H^+ + SO_4^{2-} + 2e^- \rightarrow PbSO_4 + 2H_2O$	+1.69 V	$Sn^{2+} + 2e^- \rightarrow Sn$	-0.14 V
$Au^+ + e^- \rightarrow Au$	+1.69 V	$Ni^{2+} + 2e^- \rightarrow Ni$	-0.23 V
$Pb^{4+} + 2e^- \rightarrow Pb^{2+}$	+1.67 V	$V^{3+} + e^- \rightarrow V^{2+}$	-0.26 V
$2HClO + 2H^+ + 2e^- \rightarrow Cl_2 + 2H_2O$	+1.63 V	$Co^{2+} + 2e^- \rightarrow Co$	-0.28 V
$Ce^{4+} + e^- \rightarrow Ce^{3+}$	+1.61 V	$In^{3+} + 3e^- \rightarrow In$	-0.34 V
$MnO_4^- + 8H^+ + 5e^- \rightarrow Mn^{2+} + 4H_2O$	+1.51 V	$PbSO_4 + 2e^- \rightarrow Pb + SO_4^{2-}$	-0.36 V
$Au^{3+} + 3e^- \rightarrow Au$	+1.40 V	$Cd^{2+} + 2e^- \rightarrow Cd$	-0.40 V
$Cl_2 + 2e^- \rightarrow 2Cl^-$	+1.36 V	$Cr^{3+} + e^- \rightarrow Cr^{2+}$	-0.41 V
$Cr_2O_7^{2-} + 14H^+ + 6e^- \rightarrow 2Cr^{3+} + 7H_2O$	+1.33 V	$Fe^{2+} + 2e^- \rightarrow Fe$	-0.44 V
$O_2 + 4H^+ + 2e^- \rightarrow Mn^{2+} + 2H_2O$	+1.29 V	$Zn^{2+} + 2e^- \rightarrow Zn$	-0.76 V
$MnO_2 + 4H^+ + 2e^- \rightarrow Mn^{2+} + 2H_2O$	+1.21 V	$2H_2O + 2e^- \rightarrow H_2 + 2OH^-$	-0.83 V
$Pt^{2+} + 2e^- \rightarrow Pt$	+1.20 V	$Cr^{2+} + 2e^- \rightarrow Cr$	-0.91 V
$Br_2 + 2e^- \rightarrow 2Br^-$	+1.09 V	$Mn^{2+} + 2e^- \rightarrow Mn$	-1.18 V
$2Hg_2^{2+} + 2e^- \rightarrow Hg_2^{2+}$	+0.92 V	$V^{2+} + 2e^- \rightarrow V$	-1.19 V
$ClO^- + H_2O + 2e^- \rightarrow Cl^- + 2OH^-$	+0.89 V	$ZnS + 2e^- \rightarrow Zn + S^{2-}$	-1.44 V
$Ag^+ + e^- \rightarrow Ag$	+0.80 V	$Al^{3+} + 3e^- \rightarrow Al$	-1.66 V
$Hg_2^{2+} + 2e^- \rightarrow 2Hg$	+0.79 V	$Mg^{2+} + 2e^- \rightarrow Mg$	-2.36 V
$Fe^{3+} + e^- \rightarrow Fe^{2+}$	+0.77 V	$Na^+ + e^- \rightarrow Na$	-2.71 V
$MnO_4^- + 2H_2O + 3e^- \rightarrow MnO_2 + 4OH^-$	+0.60 V	$K^+ + e^- \rightarrow K$	-2.92 V
$I_2 + 2e^- \rightarrow 2I^-$	+0.54 V	$Li^+ + e^- \rightarrow Li$	-3.05 V
$O_2 + 2H_2O + 4e^- \rightarrow 4OH^-$	+0.40 V		
$Cu^{2+} + 2e^- \rightarrow Cu$	+0.34 V	*All ions are aqueous (aq)	
$Hg_2Cl_2 + 2e^- \rightarrow 2Hg + 2Cl^-$	+0.27 V	*Most neutral species are solids (s) with some being liquids (l), gases (g) or even aqueous (aq)	
$AgCl + e^- \rightarrow Ag + Cl^-$	+0.22 V		
$NO_3^- + H_2O + 2e^- \rightarrow NO_2^- + 2OH^-$	+0.01 V		
$2H^+ + 2e^- \rightarrow H_2$	0.00 V		